

Note: - You have four choices for each objective type question as A, B, C and D. The choice which you think is correct; fill that circle in front of that question number in your answer book. Use marker or pen to fill the circles. Cutting or filling up two or more circles will result no mark.

SECTION-A

Q.1	Questions	A	B	C	D
1.	Splitting of spectral lines when atoms are subjected to strong electric field is called:	Zeeman effect	Stark effect	Photoelectric effect	Compton effect
2.	The molecules of CO_2 in dry ice form the:	Ionic crystals	Covalent crystals	Molecular crystals	Any type of crystals
3.	NH_3 shows a maximum boiling point among the hydrides of group-V elements due to:	Very small size of nitrogen	Lone pair of electrons present on nitrogen	Enhanced electronegative character of nitrogen	Pyramidal structure of NH_3
4.	The deviation of a gas from ideal behaviour is maximum at:	$-10^\circ C$ and 5 atm	$-10^\circ C$ and 2 atm	$100^\circ C$ and 2 atm	$0^\circ C$ and 2 atm
5.	Number of molecules in one dm^3 of water is close to:	$\frac{6.02}{22.4} \times 10^{23}$	$\frac{12.04}{22.4} \times 10^{23}$	$\frac{18}{22.4} \times 10^{23}$	$55.6 \times 6.02 \times 10^{23}$
6.	Which one of the given is not purified by sublimation?	Naphthalene	Silicon dioxide	I_2	Benzoic acid
7.	Solvent extraction is an equilibrium process and it is controlled by:	Law of mass action	The amount of solvent used	Distribution law	The amount of the solute
8.	27g of Al will react completely with how much mass of O_2 to produce Al_2O_3 .	8g of oxygen	16g of oxygen	32g of oxygen	24g of oxygen
9.	Many elements have fractional atomic masses. This is because:	The mass of the atom is itself fractional	Atomic masses are average masses of isobars	Atomic masses are average masses of isotopes	Atomic masses are average masses of isotopes proportional to their relative abundance
10.	Stronger the oxidizing agent, greater is the:	Oxidation potential	Reduction potential	Redox potential	E.M.F of cell
11.	The rate of reaction:	Increases as the reaction proceeds	Decreases as the reaction proceeds	Remain the same as the reaction proceeds	May decrease or increase as the reaction proceeds
12.	A solution of glucose is 10% w/v. The volume in which 1g mole of its dissolved will be:	1 dm^3	1.8 dm^3	200 cm^3	900 cm^3
13.	The pH of 10^{-3} mole dm^{-3} of an aqueous solution of H_2SO_4 is:	3.0	2.7	2.0	1.5
14.	Calorie is equivalent to:	0.4184 J	41.84 J	4.184 J	418.4 J
15.	The structure of water molecule is:	Angular	Linear	Trigonal	Trigonal pyramidal
16.	Which of the given molecules has zero dipole moment?	NH_3	$CHCl_3$	H_2O	BF_3
17.	The wave number of the light emitted by a certain source is $2 \times 10^6\text{ m}^{-1}$. The wavelength of this light will be:	500 nm	500 m	200 nm	$5 \times 10^7\text{ m}$